

Minimization of Phenol by Natural Occurrence Adsorbent

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ABSTRACT

Phenol and chlorophenols are long-lived pollutants frequently found in industrial effluents. Phenols are widely used for the commercial production of a wide variety of resins including phenolic resins, epoxy resins and adhesives, and polyamide for various applications. Adsorption process has been proven one of the best water treatment technologies around the world and the removal of diverse types of pollutants from water. However, widespread use of commercial adsorbent is sometimes restricted due to its higher costs. Attempts have been made to develop inexpensive adsorbents utilizing for the reduction of phenol from water. Four types of adsorbent clay, algae, moringa oleifera and rice husk has been used. Among all rice husk shown 97 % of phenol adsorption at 1mm particle size, pH 4, 3 g/l dosing and 150min contact time.

Keywords: Algae; Bioadsorption; Clay; Effluent; Ions

1. INTRODUCTION

Environmental pollution is an emerging threat and of great concern in today's context pertaining to its effect on the ecosystem. Water pollution is one of the greatest concerns now a day. In recent years, considerable attention has been paid to industrial wastes discharged to land and surface water [1]. Industrial effluents often contain various toxic metals, harmful dissolved gases, and several organic and inorganic compounds. These may accumulate in soil in excessive quantities in long-term use, ultimately physiologically adverse effects on crop productivity [2]. The worldwide rise in population and the industrialization during the last few decades have resulted in ecological unbalance and degradation of the natural resources [3]. One of the most essential natural resources, which have been the worst victim of population explosion and growing industrialization, is water. Huge quantity of wastewater generated from human settlement and industrial Sectors accompany the disposal system either as municipal wastewater or industrial wastewater [3]. This wastewater is enriched with varied pollutants and harmful both for human being and the aquatic flora and fauna, finds its way out into the nearly flowing or stationary water bodies and thus makes natural sources of

water seriously contaminated [5]. It has been estimated that over 5 million chemical substances produced by industries have been identified and about 12000 of these are marketed which amount to around half of the total production. Due to discharge of toxic effluents long-term consequence of exposure can cause cancer, delayed nervous damage, malformation in urban children, mutagenic changes, neurological disorders etc. [6]. Various acid manufacturing industries discharge acidic effluent, which not only make the land infertile. But make the water of the river acidic also. The high acidity causes stomach diseases and skin ailments in human beings [7].

Thus it is imperative to purify and recycle wastewater in view of reduced availability and deteriorating water quality. Phenol along with other xenobiotic compounds is one of the most common contaminants present in effluents from chemical process industries. Even at lower concentration these compounds adversely affect aquatic as well as human life [8,9]. Also these compounds form complexes with metal ions discharged from other industries, which are carcinogenic in nature. It is water soluble and highly mobile. This imparts medicinal taste and odour even at much lower concentration of 2 µg/l and it is lethal to fish at concentrations of 5-25 mg/l [10]. The maximum permitted concentration level of phenol being 0.5-1 mg/l for industrial wastewater and 1µg/l for drinking water [11,12]. So it is highly essential to save the water resources and aquatic life by removing these compounds from wastewater before disposal.

The main sources of phenolic wastewater are coal chemical plants, oil refineries, petrochemical industries, fibre glass units, explosive manufacture, phenol-based polymerization process, pharmaceuticals, plastic, paints and varnish producing units, textile units making use of organic dyes, antiseptics, antirust products, biocides, photographic chemicals and smelting and related metallurgical operations, etc. [13,14]. The conventional methods of treatment of phenolic and nitrate-nitrogen wastewater are largely physical and chemical processes but these processes led to secondary effluent problems due to formation of toxic materials such as cyanates, chlorinated phenols, hydrocarbons, etc. These methods are mainly chlorination, ozonation, solvent extraction, incineration, chemical oxidation, membrane process, coagulation, flocculation, adsorption, ion exchange, reverse osmosis, electrolysis, etc. [9-14]. The goal of study is minimized the phenol contain wastewater by low cost treatment. Natural occurrence materials are used as adsorbent for reduction. The effect of pH, dosing, contact time, isotherm has been also studied [25-32].

2. MATERIAL AND METHODS

2. 1. Material

2. 1. 1. Adsorbent

The adsorbent clay (C), rice husk (RH), algae (A) and *Moraga oleifera* (MO) was arranged from local region. All the adsorbent were washed with generous amounts of distilled water and dried in an oven at 60 °C overnight. Then, they were sieved in the size pore range from 0.25 to 1.5 mm.

2. 1. 2. Phenol water

The synthetic water was made in labotary; The phenolic compounds used in this study were phenol, 2-chlorophenol (2-CP) and 4-chlorophenol (4-CP). A stock solution was prepared by dissolving 1.0 g of phenol, 2-CP or 4-CP in 1 dm³ of deionised water. The

phenolic solutions for the sorption experiments were prepared by diluting the stock solution to give different concentrations within the range 10 to 500 mg dm⁻³ for phenol and 10 to 1000 mg dm⁻³ for chlorophenols [15].

2. 2. Methods

The batch experiments were conducted by mixing 100 mL of synthetic water prepared in the laboratory. The adsorbent collected from the local region was put for 2 hr in 100 mL glass bottles at room temperature. After 2 hr mixing, the mixtures were centrifuged at 4000 rpm for 10 minutes and decanted before final analysis. Adsorption studies were carried out by batch process. A measure amount (50 ml) of wastewater solution with measure amount of salts (or pollutant minerals) at concentration C_i was mixed with a one gram of dried and grinded adsorbent without any pre-treatment. The solutions put in contact with the adsorbents were maintained at a constant temperature of 25 °C in a water bath thermostat, the mixture being vigorously stirred by means of a magnetic stirrer. The sampled solutions were centrifuged at 5000 rpm for 15 min with a centrifugation machine. The variation of the adsorbed ions concentration (C_r) represented in the figures is defined as C_r = C₀ - C_f for the ratio 25 g/l of mass/solution, and the removal percentage of pollutant from wastewater (E) on micro-particles of clay, was calculated from equation mention below.

$$E(\%) = \frac{C_0 - C_f}{C_0} \times 100$$

where C₀ and C_f are the initial and equilibrium concentration of pollutant solution (mg/l), respectively. The capacity (mg/g) of pollutants adsorbed from wastewaters by micro-particles of clay was calculated from

$$q_a = \frac{C_0 - C_f}{m} \times v$$

where, V is the total solution volume (ml), m the weight of the materials (g).

3. RESULT AND DISCUSSION

3. 1. Effect of particle size

The effect of particle size was carried out at 2 g/l dosing, 90 min experimental time, pH 8 with 0.25 mm to 1.50 mm particle size. The result represent in Fig. 1. From the results it was found that percentage reduction of phenol was increase with increase with the particle size, after that it became constant. The maximum phenol reduction 77 % moringa oleifera (MO), 74 % algae (A), 72 % rice husk (RH) and 69 % clay (C) at 1mm particle sized respectively. The particle size play important role for adsorption due to it provides more surface area for adsorbs. At the surface of the solids, there are unbalanced forces of attraction which are responsible for adsorption. In cases where the adsorption is due to weak van der Waals forces, it is called physical adsorption. On the other hand, there may be a chemical bonding between adsorbent and adsorbate molecule and such type of adsorption is referred as chemisorption [16].

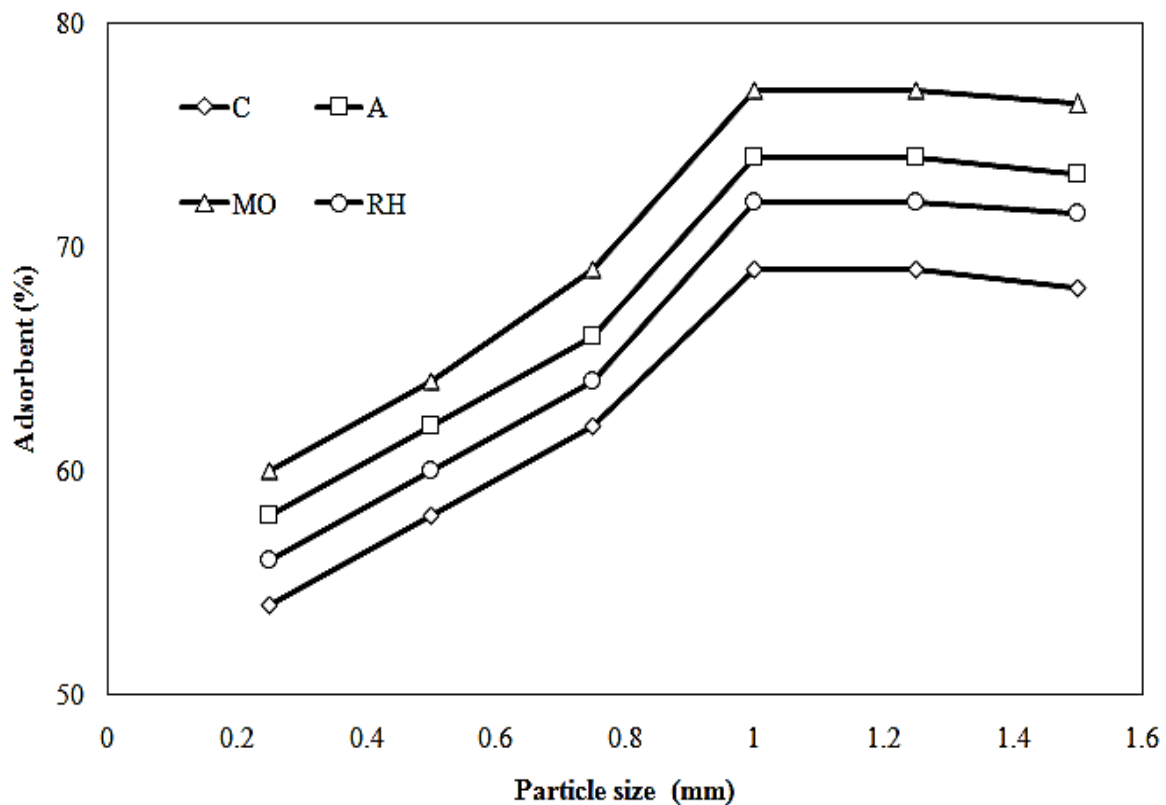


Fig. 1. Effect of particles size on adsorption of phenol.

3. 2. Effect of pH

The effect of pH was carried out at 2 g/l dosing, 1mm particles size, 90 min experimental time with different pH (2 to 12). The pH was adjusted by 0.1 M of HCL and 0.1 NaOH solutions, which is shown in Fig. 2. The result shows that maximum 85 % phenol adsorption occurred with rice husk (RH) at pH 4. The adsorbent of moringa oleifera shows 83 %, at pH 6, clay 80 % at pH 4 and algae 78 % at pH 6. The pH of the solution play an important role in the whole adsorption process and particularly on the adsorption capacity, influencing not only the surface charge of the adsorbent, the degree of ionization of the material present in the solution and the dissociation of functional groups on the active sites of the adsorbent. The hydrogen and hydroxyl ions are adsorbed quite strongly, and therefore, the adsorption of other ions is affected by the pH of the solution. It is a commonly known fact that the anions are favorably adsorbed by the adsorbent at lower pH values due to presence of H⁺ ions and at high pH values, cations are adsorbed due to the negatively charged surface sites [17]. Mahvi et al. investigated the potential of rice husk and rice husk ash for phenol removal from aqueous solution [18].

It was observed by them that rice husk ash was more effective than rice husk for phenol removal and can be used as an efficient adsorbent for the removal of phenolic compounds from wastewater. Chemically and thermally treated rice husk (RHT) was used as an adsorbent for the removal of 2,4-dichlorophenol from aqueous solution by Akhtar et al. [19].

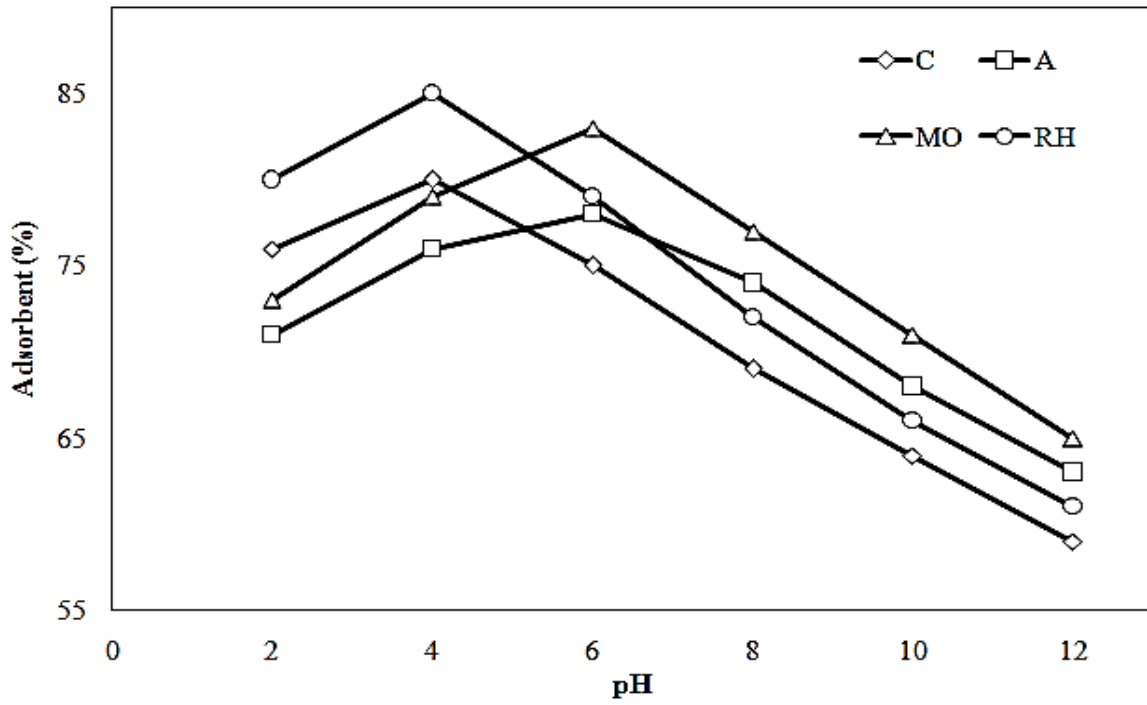


Fig. 2. Effect of pH on adsorption of phenol.

3. 3. Effect of dosing

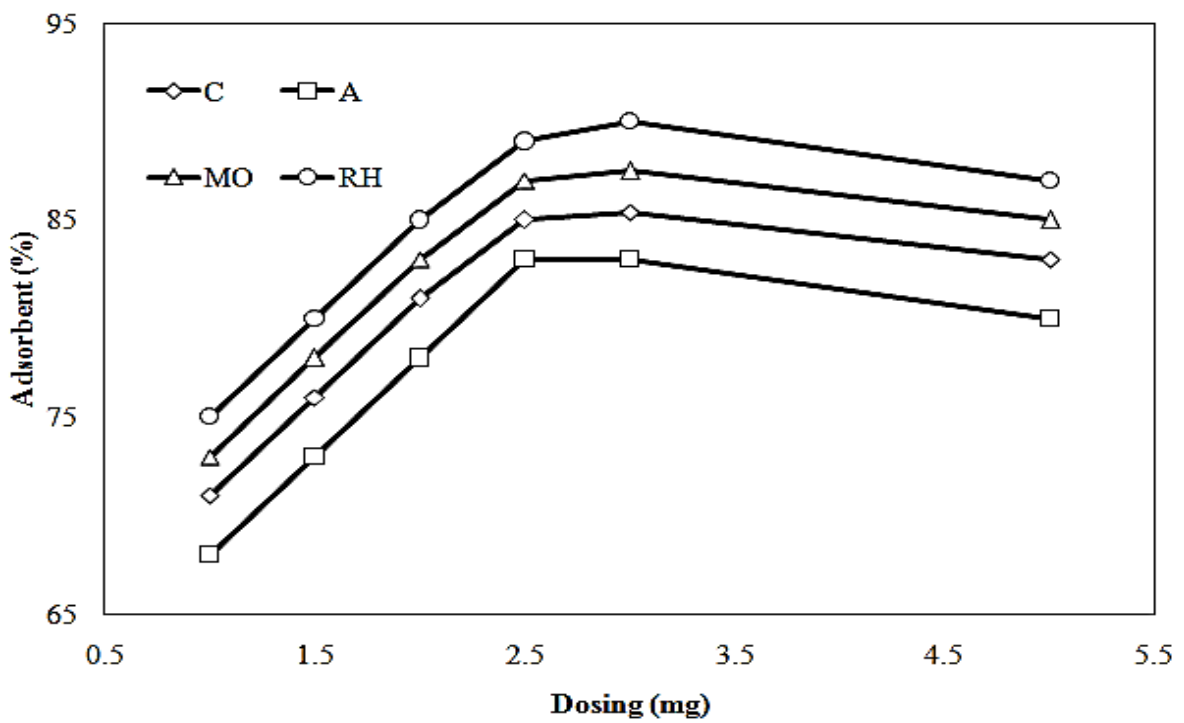


Fig. 3. Effect of dosing on phenol adsorption.

The effect of dosing was carried out at optimum pH (clay and rice husk pH and (algae and moringa oleifera at pH 6), 1 mm particle size, 90 min experimental time with dosing from 1 g/l to 5 g/l. The result represent in Fig. 3. From the result it was found that maximum 90 % adsorption shown by rice husk at 3 g/l dosing. At the same dosing of clay shows 85.4 % adsorption, moringa oleifera and algae 87 % and 83 % adsorption at 2.5 g/l dosing.

The adsorption increase with increase in dose may be due to due to the greater surface area with more functional groups consequent to the increase in the number of particles with more number of exchangeable sites for adsorption and saturation occurs as a result of non availability of exchangeable sites on the adsorbent.

The maximum sorption (98 ± 1.2 %) was achieved for RHT from 6.1×10^{-5} mol/dm³ of sorbate solution using 0.1 g of rice husk for 10 min agitation time at pH 6 and 303 K, which was comparable to commercial activated carbon (96.6 ± 1.2 %), but significantly higher than chemically treated rice husk (65 ± 1.6 %) and untreated rice husk (41 ± 2.3 %). Several other researchers also studied the application of rice husk as adsorbent for the removal of different pollutants from water [20,21].

3. 4. Effect of contact time

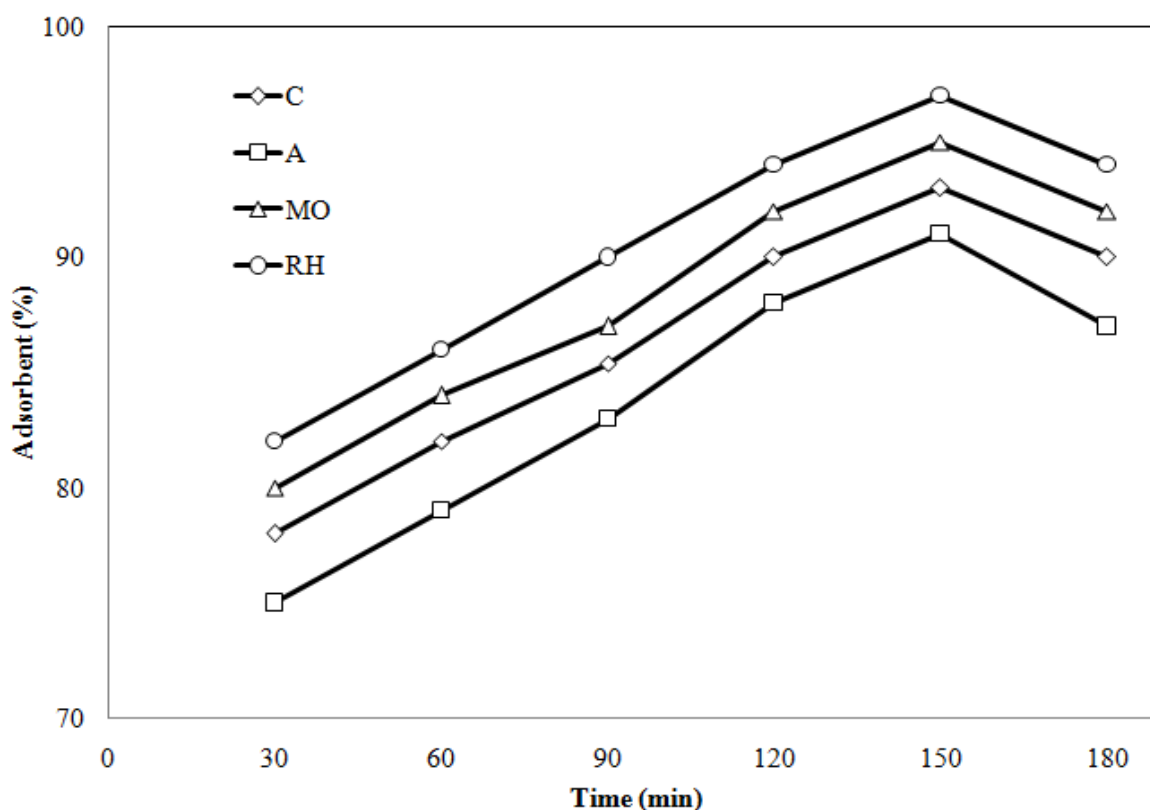


Fig. 4. Effect of contact time on phenol adsorption.

The effect of contact time was carried out optimum pH, optimum dosing, and 1mm particle size with experimental time variation from 30 min to 180 min. The result represented in Fig. 4. From the result it was found that maximum 97 % adsorption shown by rice husk at

150 min of experimental time. Almost all adsorbate show good efficiency at 150 min of contact time, moringa oleifera 95 %, clay 93 % and algae 91 % respectively.

The adsorption increase with increase with time may due to that time is sufficient to reach equilibrium. After reaching maximum adsorptions it was decrease may this is because the agitation of the solution facilitates the rate of transport of the adsorbate species from the outer sites to the interior site of the adsorbent. As the vacant sites are occupied the rate of adsorption becomes decrease [22].

3. 5. Isotherm Study

Adsorption isotherm helps in determining the properties of the adsorbents such as pore volume pore size or energy distribution and specific surface area. The isotherm curve can also be utilized to obtain information concerning the desorption mechanism strictly connected with interaction between the adsorbent and adsorbate molecules. Therefore, the efficiency of an industrial adsorbent can be accessed through this curve.

The correct interpretation of experimental adsorption isotherm can be realized in terms of some mathematical equations called adsorption isotherm model equations [23]. Such equations are derived assuming an ideal physical model for the adsorption system. The model assumptions are usually a result of experimental observations. For this study Langmuir model was assess.

The Langmuir adsorption model assumes monolayer adsorption (the adsorbed layer is one molecule in thickness), with adsorption can only occur at a finite (fixed) number of definite localized sites, that are identical and equivalent, with no lateral interaction and strict hindrance between the adsorbed molecules, even on adjacent sites. In its derivation, Langmuir isotherm refers to homogeneous adsorption, which each molecule possess constant enthalpies and sorption activation energy (all sites possess equal affinity for the adsorbate), with no transmigration of the adsorbate in the plane of the surface.

Moreover, Langmuir theory stipulates a rapid decrease of the intermolecular attractive forces to the rise of distance [24].

The equation proposed by Langmuir was universally applicable to chemisorption with some restrictions involving physical adsorption where it is applicable adsorption on solid surface with one type of adsorption active center.

The Langmuir equation is given by:

$$\frac{C_e}{q_e} = \frac{1}{Q_0 b} + \frac{C_e}{Q_0}$$

where 'C_e' is the equilibrium concentration and 'q_e' is the amount of adsorbate adsorbed per gram of adsorbent at equilibrium (mg/g); 'Q₀' and 'b' are Langmuir constants related to the sorption capacity and intensity respectively.

The notations q_e and C_e are the concentration of adsorbate in solid phase at equilibrium (mg/g) and the concentration of adsorbate in liquid phase at equilibrium (mg/L) respectively. A linear plot was obtained when C_e/q_e was plotted against C_e over the entire concentration range evaluated.

The straight line indicates that the adsorption complies with the Langmuir mode as shown on Fig. 5.

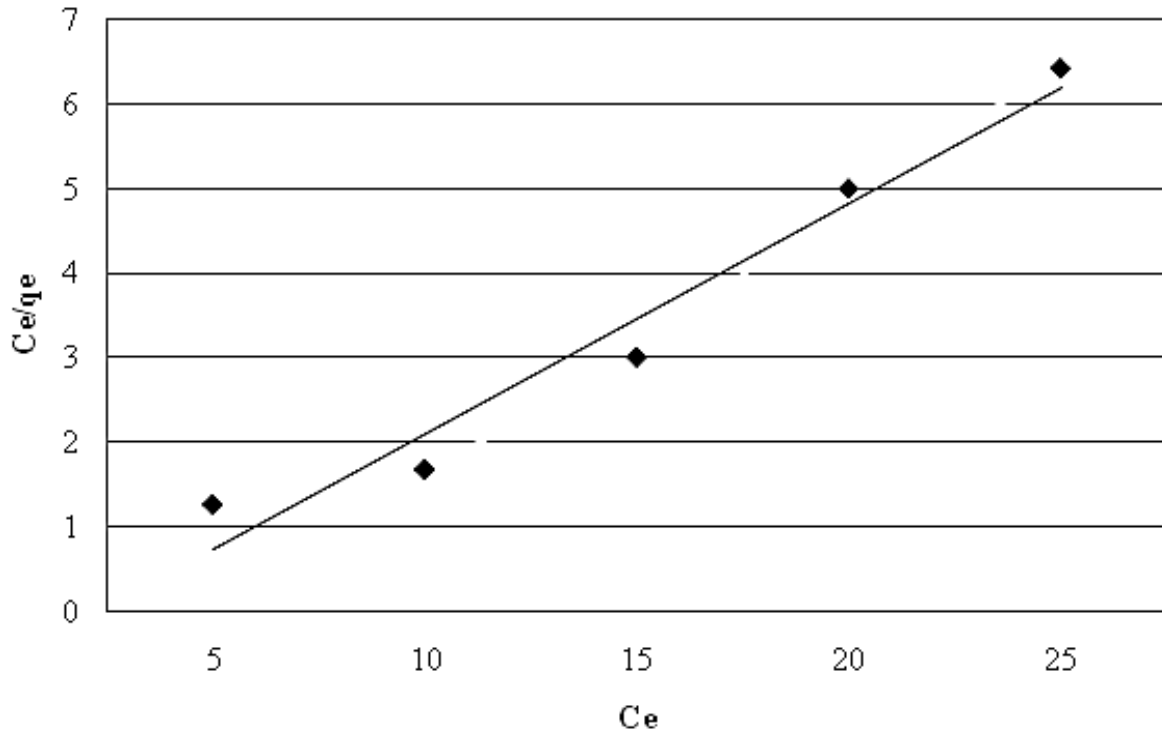


Fig. 5. Langmuir adsorption isotherm for rice husk.



Fig. 6. Rice husk (a) raw (b) grinded (c) surface of rice husk.

In Fig. 6(a) the rice husk which is collected from rice industry found brownish yellow in color. That can be dried at 60 °C in oven for around 12 h. Dry rice husk was grinded with the help of grinder and make it fine with the help of mesh according to requirement, which is shown in Fig. 6(b).

Before adsorption the rice husk was scanning with electron micrographic equipment which is shown in Fig. 6(c). It was found rice husk greater surface area which is responsible for high efficiency.

4. CONCLUSION

It is concluding that selection and identification of an appropriate low-cost adsorbent is one of the key issues to achieve the maximum removal/adsorption of specific type of pollutant depending upon the adsorbent-adsorbate characteristics. By comparatively study between clay, algae, moringa oleifera and rice husk for adsorption of phenol rice husk shown 97 % adsorption at 150 min of contact time. Other adsorbent moringa oleifera 95 %, clay 93 % and algae 91 % adsorption shows at same condition. Rice husk are available easy and economical effectiveness for the reduction of phenol from water. The disposal of rice husk after uses was ecofriendly for soil.

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